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range for E_T . Where the surface is bare soil, then the K_c will be quite similar to the K_c ini predicted in Chapter

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Set 3. 0.1M barium chloride,
BaCl₂. 0.1M magnesium nitrate,

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$\text{Mg}(\text{NO}_3)_2$. 0.2M sodium

chromate, Na_2CrO_4 . 0.2M

aluminum sulfate, $\text{Al}_2(\text{SO}_4)_3$.

0.2M potassium chromate,

K_2CrO_4 . 0.2M silver nitrate,

AgNO_3 . SAFETY FIRST! In this lab,
the solutions you use may contain
harmful materials.

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The aim of this part of the experiment is to determine the factors that control whether or not a precipitate forms, i.e. is it the sum of the ion concentrations (in this case $[Ag^+] + [Cl^-]$) or the product of the ion concentrations (in this

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Aqueous silver nitrate (AgNO_3) is added to a solution containing potassium chloride (KCl), and the precipitation of a white solid, silver chloride (AgCl), is observed:

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$\text{AgNO}_3 (\text{aq}) + \text{KCl} (\text{aq}) \rightarrow \text{AgCl} (\text{s}) + \text{KNO}_3 (\text{aq})$ Note that the product silver chloride is the precipitate, and it is designated as a solid.

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overview of downstream

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processing. After completing this chapter, students will be able to: describe the general elements of the downstream purification process for a typical monoclonal antibody-based therapeutic, starting with the bioreactor harvest pool and finishing with

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